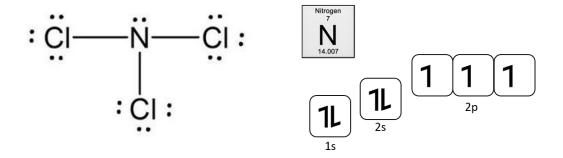
Hybridisation Practice

Determine the hybridisation of each central atom:

Recall that Pi bonds are made from unhybridised orbitals, and that the 3d orbital is used for expanded octets (not the 4s).

$$H - \ddot{O} : \\ \ddot{S} \\ \ddot{S} \\ \ddot{O} : \ddot{O} : \\ \ddot{O} : \ddot{O}$$



$$\begin{array}{c|c}
F & N \\
N & IL \\
1s & 1s
\end{array}$$